

Question Bank

1 Mark Questions

- 1) Write the structure of N-Methylethanamine.
- 2) Write the structure of Prop-2-en-1-amine
- 3) Why do amines behave as nucleophiles?
- 4) Why is an alkylamine more basic than ammonia?
- 5) Arrange the following compounds in increasing order of basic strength in their aqueous solutions: NH_3 , CH_3NH_2 , $(\text{CH}_3)_3\text{NH}$, $(\text{CH}_3)_2\text{NH}$.
- 6) Why is aniline soluble in aqueous HCl?
- 7) What is the role of pyridine in the acylation reaction of amines?
- 8) Why is benzene diazonium chloride not stored and is used immediately after its preparation?
- 9) Arrange the following in the increasing order of their solubility in water:
 $\text{C}_6\text{H}_5\text{NH}_2$, $(\text{C}_2\text{H}_5)_2\text{NH}$, $\text{C}_2\text{H}_5\text{NH}_2$.
- 10) Arrange the following compounds in increasing order of basic strength:
 $\text{C}_6\text{H}_5\text{NH}_2$, $\text{C}_6\text{H}_5\text{N}(\text{CH}_3)_2$, $(\text{C}_6\text{H}_5)_2\text{NH}$, CH_3NH_2
- 11) Give the IUPAC name of $\text{H}_2\text{N}-\text{CH}_2-\text{CH}_2-\text{CH}=\text{CH}_2$
- 12) Name one reagent used for the separation of primary, secondary & tertiary amine.

Answers of 1 Mark Questions

- 1) $\text{CH}_3\text{NHCH}_2\text{CH}_3$
- 2) $\text{H}_2\text{NCH}_2\text{CH}=\text{CH}_2$
- 3) Due to the presence of lone pair of electrons on N atom
- 4) Due to +I effect of alkyl group which increases the electron density on N atom.
- 5) $\text{NH}_3 < (\text{CH}_3)_3\text{NH} < \text{CH}_3\text{NH}_2 < (\text{CH}_3)_2\text{NH}$
- 6) Because of the formation of anilinium ion which is soluble in water.
- 7) To remove the side product, i.e., HCl from the reaction mixture.
- 8) Because it is unstable in nature.
- 9) $\text{C}_6\text{H}_5\text{NH}_2 < (\text{C}_2\text{H}_5)_2\text{NH} < \text{C}_2\text{H}_5\text{NH}_2$ (Hydrogen bonding)
- 10) $(\text{C}_6\text{H}_5)_2\text{NH} < \text{C}_6\text{H}_5\text{NH}_2 < \text{C}_6\text{H}_5\text{N}(\text{CH}_3)_2 < \text{CH}_3\text{NH}_2$ (+I effect of alkyl group and -I effect of Phenyl)
- 11) But-3-en-1-amine
- 12) Hinsberg reagent (benzene sulphonyl chloride)

2 Marks Questions

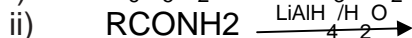
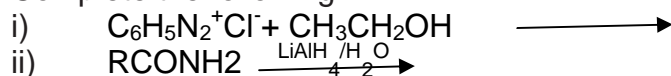
- 1) Give the chemical tests to distinguish between the following pairs of compounds:
 - i) Ethylamine and Aniline
 - ii) Aniline and Benzylamine
- 2) Identify A and B in each of the following:
 - i) $\text{CH}_3\text{CH}_2\text{Cl} \xrightarrow{\text{NaCN}} \text{A} \xrightarrow[\text{2}]{\text{Ni/H}_2} \text{B}$
 - ii) $\text{C}_6\text{H}_5\text{NH}_2 \xrightarrow[\text{-2}]{\text{NaNO}_2/\text{HCl}} \text{A} \xrightarrow[\text{6 5 2}]{\text{C}_6\text{H}_5\text{NH}_2/\text{OH}^-} \text{B}$
- 3) Give the chemical tests to distinguish between the following pairs of compounds:
 - i) Methylamine and Dimethylamine
 - ii) Aniline and N-Methylaniline
- 4) Describe the following giving the relevant chemical equation in each case:
 - i) Carbylamine reaction
 - ii) Hoffmann's bromamide reaction

5) Illustrate the following reactions giving equation in each case:

i) Gabriel phthalimide synthesis

ii) Coupling reaction

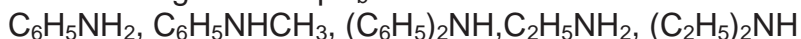
6) Complete the following :



7) In the following cases rearrange the compounds as directed:

i) In decreasing order of basic strength: Aniline, p-nitroaniline, p-toluidine

ii) In increasing order of pK_b values:



8) Explain by chemical equations :

i) Gattermann reaction

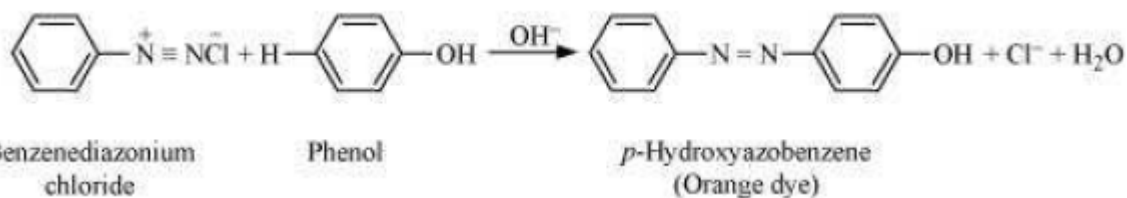
ii) Sandmeyer reaction

9) Show the mechanism of acylation of ethylamine and write the IUPAC name of the product formed.

Answers of 2 Marks Questions

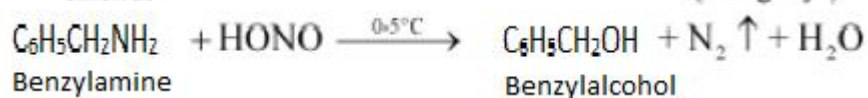
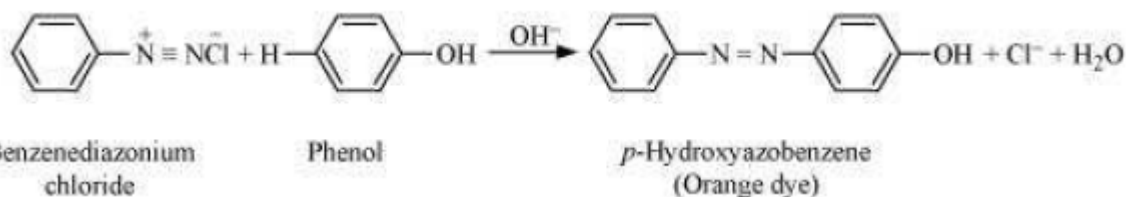
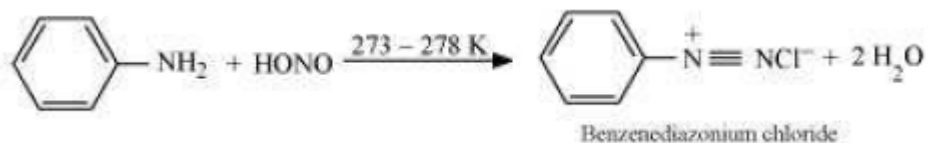
1) i) These can be distinguished by azodye test.

Aniline forms azo dye while ethylamine does not form dye.



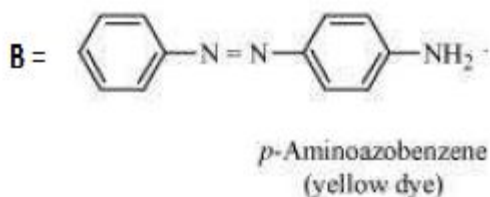
ii) These can be distinguished by azodye test.

Aniline forms azo dye while benzylamine does not form dye



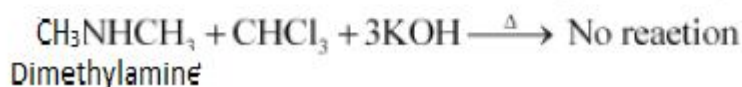
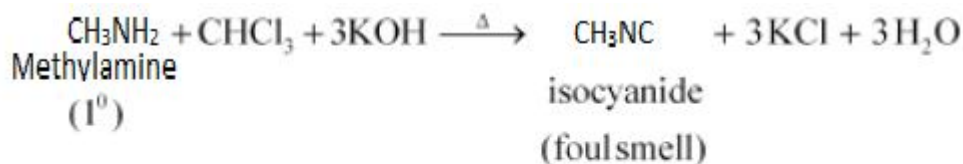
2) i) A = CH_3CH_2CN , B = $CH_3CH_2CH_2NH_2$

ii) A = $C_6H_5N_2^+Cl^-$,



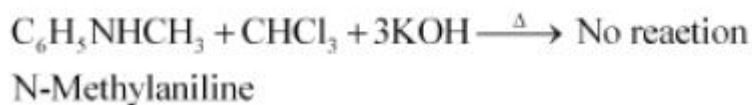
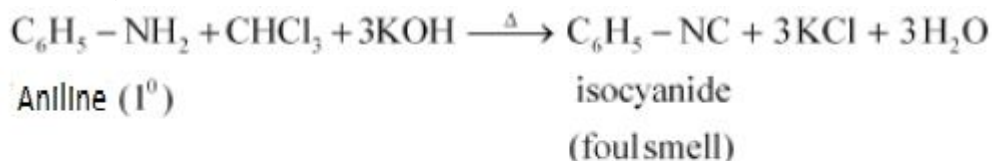
3) i) These can be distinguished by carbylamine test.

Methylamine gives this test whereas Dimethylamine does not.

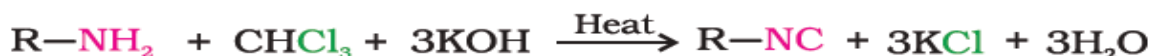


ii) These can be distinguished by carbylamine test.

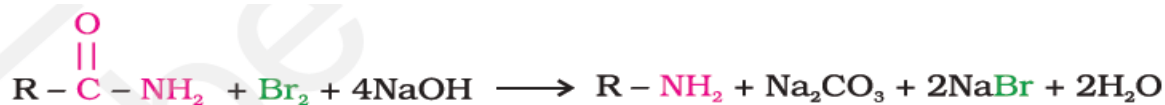
Aniline gives this test whereas N-Methylaniline does not.



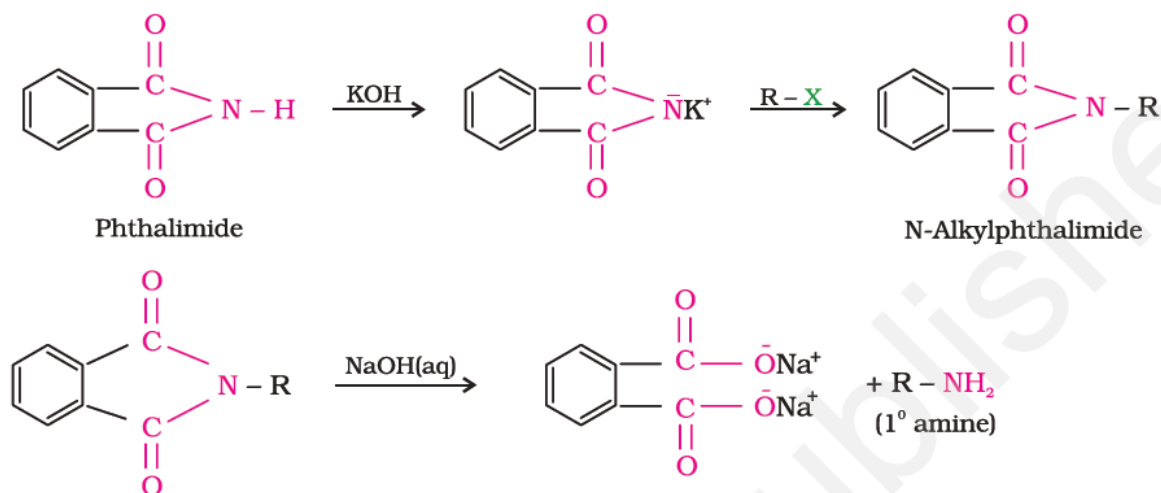
4) Carbylamine Reaction: Aliphatic and aromatic primary amines on heating with chloroform and ethanolic potassium hydroxide form isocyanides or carbylamines, having unpleasant smell.



ii) Hoffmann Bromamide Reaction: This is a method for preparation of primary amines by treating an amide with bromine in an aqueous or ethanolic solution of sodium hydroxide.



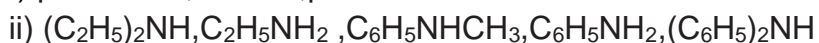
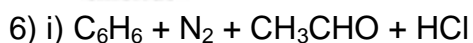
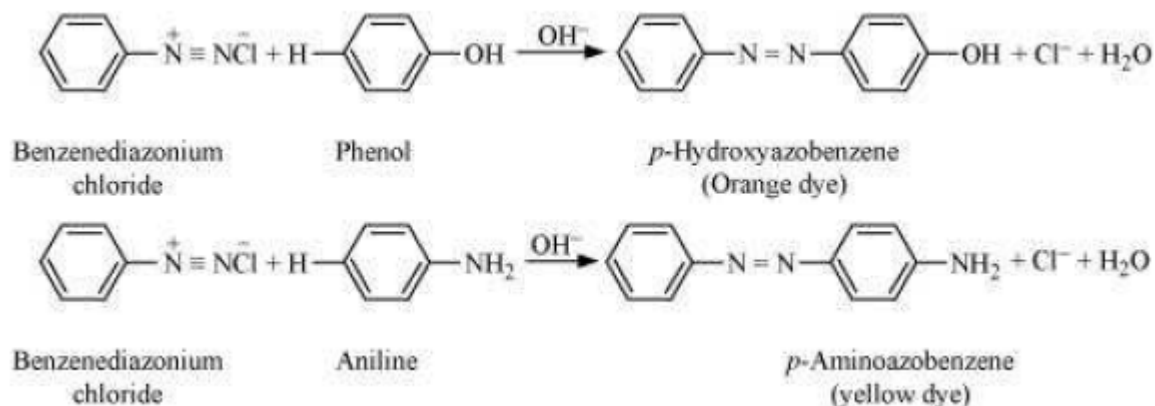
5) i) Gabriel Phthalimide Reaction: This method is used for the preparation of primary amines from phthalimide.



ii) Coupling Reaction:

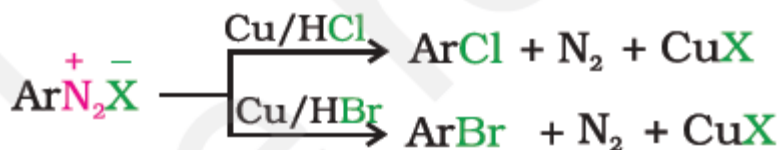
The reaction of joining two aromatic rings through -N=N- is known as coupling reaction.

Arenediazonium salts react with phenols or aromatic amines to form coloured azo compounds.



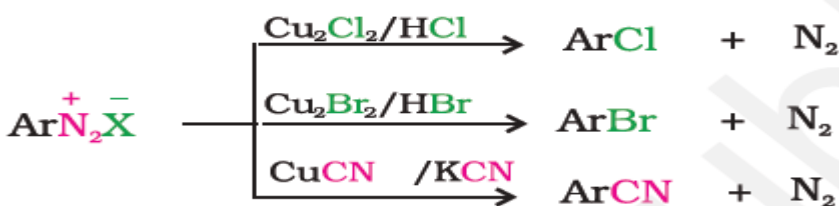
8) i) Gatterman reaction:

In this reaction, chlorine or bromine or cyanide can also be introduced in the benzene ring by treating the diazonium salt solution with corresponding halogen acid in the presence of copper powder.

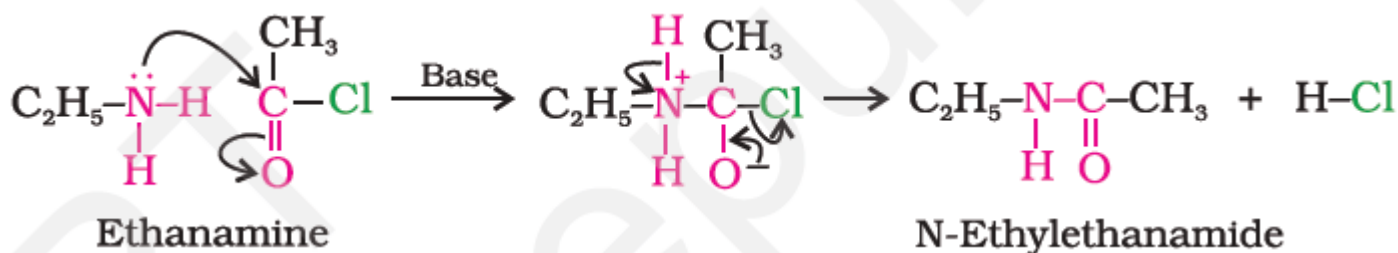


ii) Sandmeyer reaction:

In this reaction, chlorine or bromine or cyanide can also be introduced in the benzene ring by treating the diazonium salt solution with corresponding halogen acid in the presence of copper(I) ion.



9)

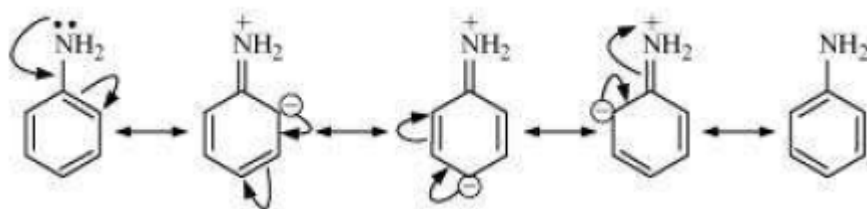


3 Marks Questions

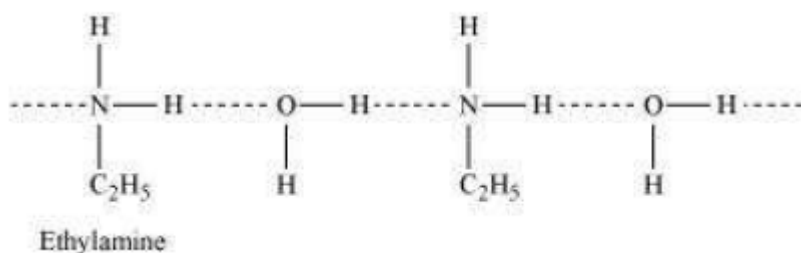
- 1) Complete the following reactions:
 - i) $\text{CH}_3\text{CH}_2\text{NH}_2 + \text{CHCl}_3 + \text{alc. KOH} \longrightarrow$
 - ii) $\text{C}_6\text{H}_5\text{N}_2^+\text{Cl}^- \xrightarrow[\text{2}]{\text{H}_2\text{O (Room Temp.)}}$
 - iii) $\text{C}_6\text{H}_5\text{NH}_2 + \text{HCl (aq)} \longrightarrow$
- 2) Write the main products of the following reactions:
 - i) $\text{C}_6\text{H}_5\text{N}_2^+\text{Cl}^- \xrightarrow[\text{2}]{\text{H}_3\text{PO}_4 + \text{H}_2\text{O}}$
 - ii) $\text{C}_6\text{H}_5\text{NH}_2 + \text{Br}_2 \text{ (aq)} \longrightarrow$
 - iii) $\text{CH}_3\text{CONH}_2 \xrightarrow[\text{2}]{\text{Br}_2 + \text{NaOH}}$
- 3) Write the main products of the following reactions:
 - i) $\text{CH}_3\text{CH}_2\text{NH}_2 \xrightarrow[\text{2}]{\text{HNO}_2, 0^\circ\text{C}}$
 - ii) $\text{C}_6\text{H}_5\text{NH}_2 + \text{CH}_3\text{COCl} \longrightarrow$
 - iii) $\text{CH}_3\text{CH}_2\text{NH}_2 + \text{C}_6\text{H}_5\text{SO}_2\text{Cl} \longrightarrow$
- 4) Account for the following:
 - (i) pK_b of aniline is more than that of methylamine.
 - (ii) Ethylamine is soluble in water whereas aniline is not.
 - (iii) Aniline does not undergo Friedel-Craft's reaction.
- 5) State reasons for the following:
 - i) Primary amines have higher boiling points than tertiary amines.
 - ii) Although $-\text{NH}_2$ group is an ortho and para directing group, nitration of aniline gives along with ortho and para derivatives meta-derivative also.
 - iii) Aromatic primary amines cannot be prepared by Gabriel phthalimide synthesis.
- 6) How are the following conversions carried out?
 - a) Aniline to nitrobenzene
 - b) Ethanamine to N-Ethylethanamide
 - c) Chloroethane to propan-1-amine
- 7) How are the following conversions carried out?
 - a) Nitrobenzene to benzoic acid
 - b) Benzene to m-bromophenol
 - c) Benzoic acid to aniline
- 8) Accomplish the following conversions:
 - i) Aniline to 2,4,6-tribromofluorobenzene
 - ii) Benzyl chloride to 2-phenylethanamine
 - iii) Chlorobenzene to p-chloroaniline

Answers Of 3 Marks Questions

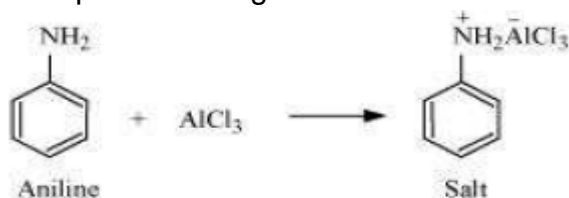
- $\text{CH}_3\text{CH}_2\text{NC} + \text{KCl} + \text{H}_2\text{O}$
 - $\text{C}_6\text{H}_5\text{OH} + \text{N}_2 + \text{HCl}$
 - $\text{C}_6\text{H}_5\text{NH}_3^+\text{Cl}^-$
- $\text{C}_6\text{H}_6 + \text{N}_2 + \text{H}_3\text{PO}_3 + \text{HCl}$
 - 2,4,6-Tribromoaniline
 - $\text{CH}_3\text{NH}_2 + \text{Na}_2\text{CO}_3 + \text{NaBr} + \text{HCl}$
- $\text{CH}_3\text{CH}_2\text{OH} + \text{N}_2 + \text{H}_2\text{O}$
 - $\text{C}_6\text{H}_5\text{NHCOCH}_3 + \text{HCl}$
 - $\text{CH}_3\text{CH}_2\text{NHSO}_2\text{C}_6\text{H}_5 + \text{HCl}$
- In aniline, lone pair of electrons on N atom is delocalised over the benzene due to which electron density on the nitrogen decreases. On the other hand, electron density on N atom increases in case of methylamine due to +I effect of methyl group. So, aniline is less basic than methylamine and hence pK_b value of aniline is more than that of methylamine.



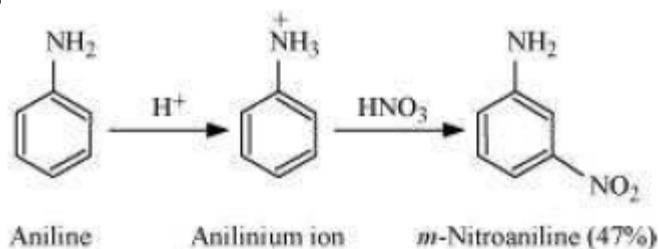
- Ethylamine is soluble in water due to intermolecular hydrogen bonding which is very very less in case of aniline due to large hydrocarbon part in aniline.



- Because aniline being Lewis base reacts with AlCl_3 , Lewis acid to form salt. Due to this N of aniline acquires + charge and hence acts as strong deactivating group for further reaction.

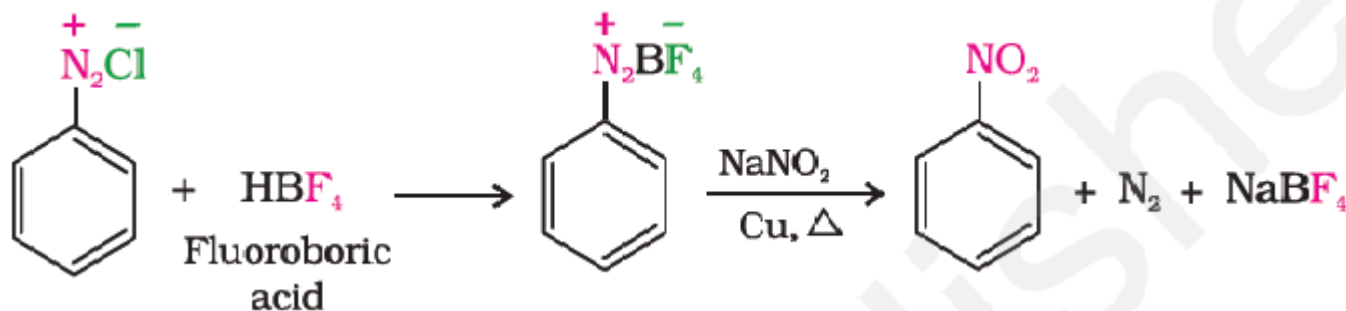
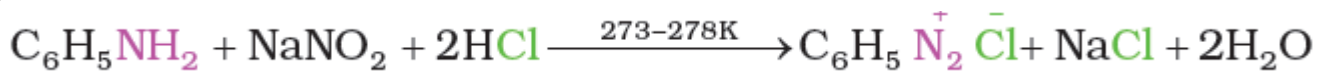


- Due to extensive intermolecular H-bonding in case of primary amines which is absent in tertiary amines
 - Because nitration is carried in acidic conditions and in acidic conditions aniline gets protonated and is converted to anilinium ion which is meta directing group.

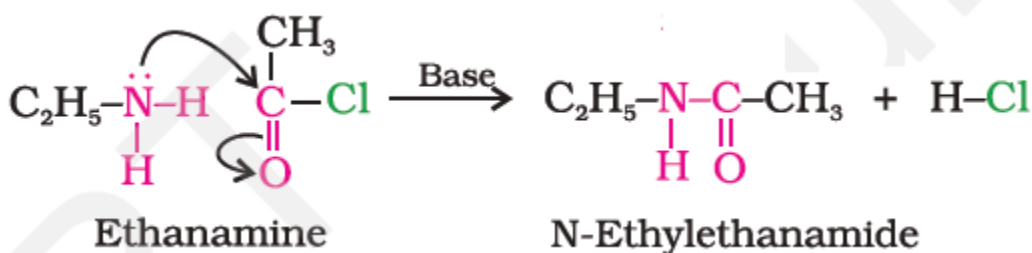


iii) Gabriel phthalimide synthesis is used for the preparation of aliphatic primary amines. It involves nucleophilic substitution (S_N2) of alkyl halides by the anion formed by the phthalimide. But aryl halides do not undergo nucleophilic substitution with the anion formed by the phthalimide. Hence, aromatic primary amines cannot be prepared by this process.

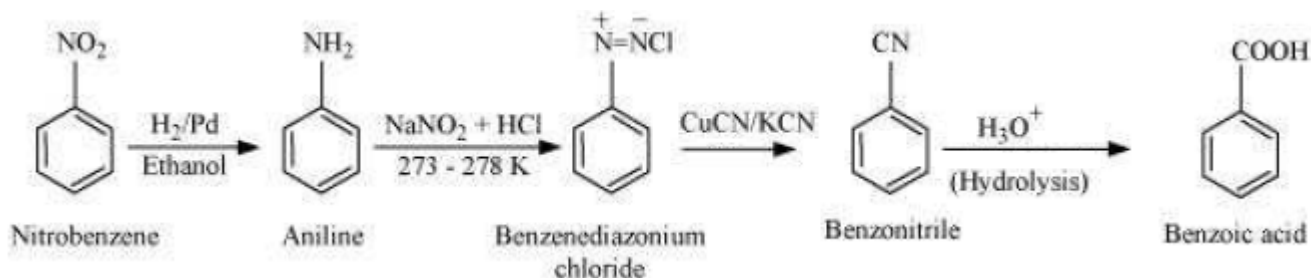
6) i)



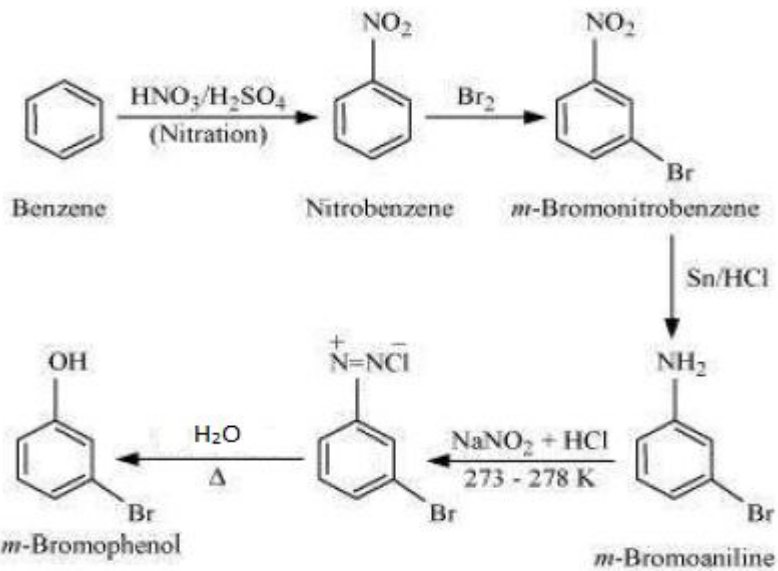
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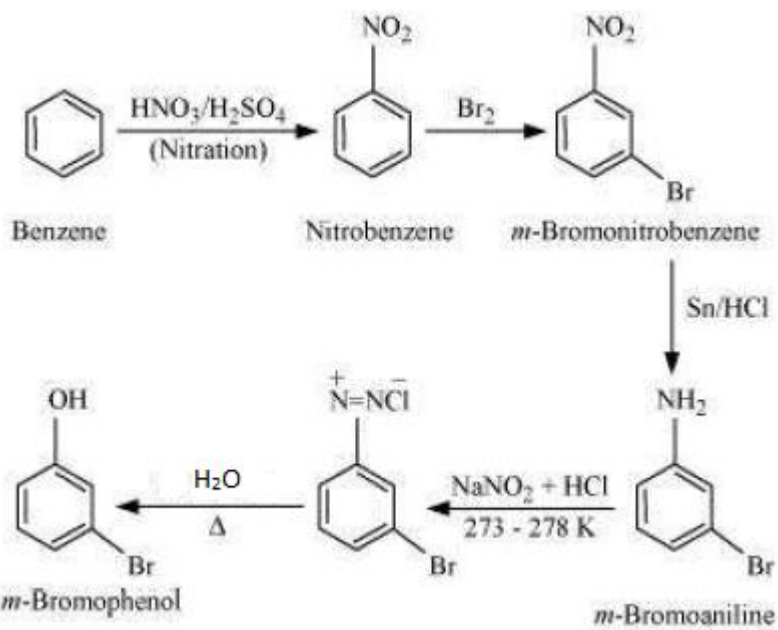
7) i)



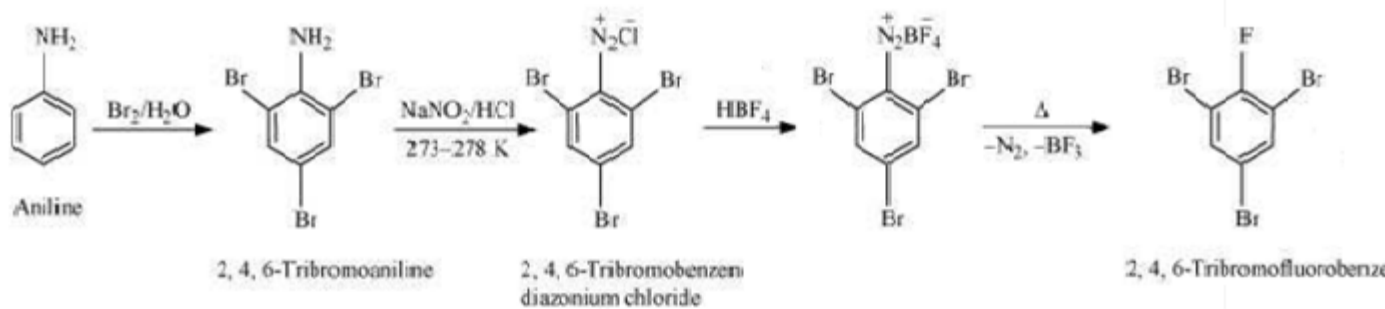
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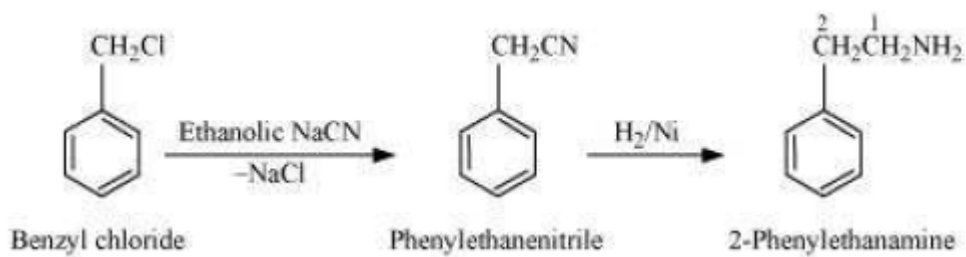
iii)



8) i)



ii)



iii)

